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Solution From the first law of thermodynamics applied for a closed system, the heat transferred is given by: $Q = \Delta U$ The change in internal energy is defined by Equation 8.3-6 of Felder & Rousseau: $\Delta U = \int C_p(T) dT$ Substituting the heat capacity equation into the integral: $\Delta U = \int_{T_1}^{T_2} C_p(T) dT$

Chapter 8 Balances on Nonreactive Processes Date:

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CHAPTER TWO - College of Engineering and Applied Science

This solutions manual contains solutions for all odd numbered problems plus a large number of solutions for even numbered problems. Of the 624 exercises in Statistical Inference, Second Edition, this manual gives solutions for 484 (78%) of them. There is an obtuse pattern as to which solutions were included in this manual.

